

FACULTY**Pharm D (6–YDC) I – Year (Main & Backlog) Examination, July 2018****Subject: Pharmaceutical Inorganic Chemistry****Time: 3 Hours****Max.Marks: 70****Note: Answer all questions from Part – A. Any Five questions from Part – B.****PART – A (10x2 = 20 Marks)**

- 1 Explain the indicators in complexometric titrations.
- 2 Explain the role of solvents in limit test for iron.
- 3 Mention the method of preparation of nitrous oxide.
- 4 What are the uses of magnesium stearate?
- 5 Mention the units of measurement of radioactivity.
- 6 Calculate the normality for 500 ml solution containing 4 gm of sodium hydroxide.
- 7 Define an error. What are the different types of errors?
- 8 Give examples for mixed and universal indicators.
- 9 Define Mohrs method.
- 10 Write about electrolyte replenishes.

PART – B (5x10 = 50 Marks)

- 11 Explain in detail about the neutralization curve for the following titrations with calculation of equivalence point and pH.
 - a) Strong acid – Strong base 5
 - b) Weak acid – Weak base 5
- 12 a) Name the magnesium compounds used as antacids. Describe the preparation, properties, assay and uses of milk of magnesia. 5
- b) Name the different types of acidifiers and give their examples. 5
- 13 a) How is end point detected in redox titrations? 5
- b) Mention pharmaceutical applications of gravimetry. 5
- 14 What are essential trace elements? Write the physiological role of copper and iodine. 10
- 15 Define limit test. Write about the principle and procedure involved in the limit test of arsenic with neat diagram. 10
- 16 Write the preparation, properties, assay and uses of sodium chloride in replacement therapy. 10
- 17 What are radiopharmaceuticals? Write about its clinical applications. 10
- 18 Explain the mechanism of action of anti-microbial agents. Give a brief account on hydrogen peroxide. 10
