

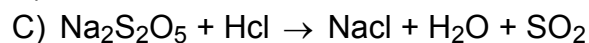
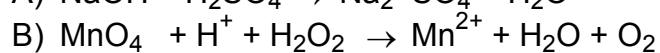
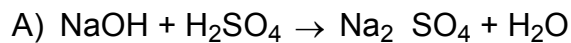
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FACULTY**B. Pharmacy 2/4 I Semester (Suppl) Examination, April / May 2018****Subject: Pharmaceutical Analysis – I (Chemical Analysis)****Time: 3 Hrs****Max. Marks: 70****Note: Answer All Questions. All Questions carry Equal Marks.**

1. a) What is error? Discuss the sources of errors and suggest the remedial measures for their control. 14
- OR**
- b) i) Define the following. 4
A) Titration B) Titrant C) Analant D) Equivalence Point
- ii) What is Primary Standard and secondary standard? Give examples. Write the ideal requirements of Primary Standard 10
- 2) a) i) Explain the following: 8
A) Common ion effect B) Solubility Product
- ii) Write a note on theories of indicators used in neutralization titration. 6
- OR**
- b) i) Discuss the modern theories of acids and bases 8
- ii) Explain the pharmaceutical importance of buffer solutions 6
- 3) a) i) Explain the principle and various steps involved in volumetric analysis 8
- ii) How do you prepare and standardize 0.1N NaOH 6
- OR**
- b) i) Write the Principle and Procedure of redox titration involving use of $K_2Cr_2O_7$ as titrant. 8
- ii) Explain the following: 6
A) Co-Precipitation B) Post – Precipitation
- 4) a) i) Discuss the masking and demasking agents with their applications in complexometry. 8
- ii) Write a note on solvents used in non-aqueous titration. 6
- OR**
- b) i) Explain the basic principles of complexometric analysis. Write the procedure for preparation and standardization of 0.1M EDTA 3+6
- ii) Write the applications of non-aqueous titrations in pharmaceutical analysis 5
- 5) a) i) Define the terms : 4
A) Molecular formula B) Empirical formula
- ii) Balance the following equations 6

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iii) Calculate the number of moles in each of the followin: 4

A) 392 m of Sulphuric acid.

B) 9.0 m of Aluminium

OR

b) i) Write mass balance expressions for a 0.0100M Solution of HCl that is in equilibrium with an excess of solid BaSO_4 8

ii) Find the percent composition of each element in water. 6

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