

Rajiv Gandhi University of Health Sciences, Karnataka

I Year B.Pharm Degree Examination – JAN-2019

Time: Three Hours**Max. Marks: 70 Marks**

PHARMACEUTICAL INORGANIC CHEMISTRY (Revised Scheme 3)

Q.P. CODE: 2605

Your answers should be specific to the questions asked
Draw neat labeled diagrams wherever necessary

LONG ESSAYS (Answer any Two)**2 x 10 = 20 Marks**

1. Discuss in detail the limit test for Iron and Sulphate.
2. Define and classify antacid and give the method of preparation and assay for magnesium hydroxide mixture.
3. Write the principle behind complexometric titration. Explain preparation and assay for calcium gluconate.

SHORT ESSAYS (Answer any Six)**6 x 5 = 30 Marks**

4. Name the effect of impurities, on the properties of pharmaceutical substances.
5. Discuss the important functions of sodium ions in the body.
6. Write the preparation, properties, assay and category for copper sulphate.
7. Explain the solvents used in nonaqueous titration with examples.
8. What is precipitation titration? Explain the assay of NaCl by Volhards method.
9. What are dental products? Add a note on role of fluoride in preventing the dental caries.
10. Write the preparation, properties, storage, labeling, and category for any two medicinal gases.
11. Write the preparation and standardization of 0.1N Hydrochloric acid.

SHORT ANSWERS**10 x 2 = 20 Marks**

12. Why nitric acid is used in the limit test for chloride?
13. Define the term astringent and bactericidal.
14. What are expectorants?
15. Name the indicators used in redox titration.
16. Define accuracy and precision.
17. What is test for purity?
18. Give the examples of protective and adsorbents.
19. Why is simethicone added in antacid preparation?
20. Define normality and equivalent weight.
21. How boric acid is prepared.
