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## Sample questions Physics

S. No  Questions  A carnot engine working between 300K has a work output of 800J per cycle. The amoun of heat energy supplied to the engine in each cycle is a) 800J b) 1600J c) 3200J d) 6400J  For hydrogen gas C <sub>P</sub> -C <sub>v</sub> =a and for oxygen gas C <sub>P</sub> -C <sub>v</sub> =b. The relation between a and b is a) a=16b b) a=b16 c) a=4b d) a=b  The equation of state corresponding to 8 g of O <sub>2</sub> is (assume O <sub>2</sub> to be an ideal gas) a) PV=8RT b) PV=8RT d) PV=RT/4 c) PV=RT/3 d)		Sample questions Physics		_	
heat energy supplied to the engine in each cycle is a) 800J b) 1600J c) 3200J d) 6400J  2. For hydrogen gas $C_P$ - $C_V$ - $a$ and for oxygen gas $C_P$ - $C_V$ - $b$ . The relation between $a$ and $b$ is a) $a$ - $16b$ b) $a$ - $b$ / $16$ c) $a$ - $a$ - $b$ d) $a$ - $b$ 3. The equation of state corresponding to $a$ g of $a$ g (assume $a$ ) to be an ideal gas) a) $a$ -	S. No	Questions	Answer	Marks	Chapter
a) 8007 b) 1600J c) 3200J d) 6400J  2. For hydrogen gas $C_P$ - $C_V$ = $a$ and for oxygen gas $C_P$ - $C_V$ = $b$ . The relation between $a$ and $b$ is a) $a$ = 16 $b$ b) $a$ = $b$ /16 c) $a$ = $b$ /16 d) $a$ = $b$ 3. The equation of state corresponding to 8 g of $O_2$ is (assume $O_2$ to be an ideal gas) a) $PV$ - $RT$ d) $PV$	1.	A carnot engine working between 300K has a work output of 800J per cycle. The amoun of	В	1	
b) 1600J c) 3200J d) 6400J  2. For hydrogen gas $C_P$ - $C_v$ - $a$ and for oxygen gas $C_P$ - $C_v$ - $b$ . The relation between $a$ and $b$ is a) $a$ - $16b$ b) $a$ - $b$ / $16$ c) $a$ - $b$ / $16$ c) $a$ - $a$ - $b$ d) $a$ - $b$ 3. The equation of state corresponding to $8$ g of $O_2$ is (assume $O_2$ to be an ideal gas) a) $PV$ - $RT$ b) $PV$ - $RT$ / $4$ c) $PV$ - $RT$ / $4$ c) $PV$ - $RT$ / $4$ d) $PV$ -		heat energy supplied to the engine in each cycle is			
c) 3200J d) 6400J  2. For hydrogen gas $C_P$ - $C_V$ = $a$ and for oxygen gas $C_P$ - $C_V$ = $b$ . The relation between $a$ and $b$ is a) $a=16b$ b) $a=b/16$ c) $a=4b$ d) $a=b$ 3. The equation of state corresponding to 8 g of $O_2$ is (assume $O_2$ to be an ideal gas) a) $PV$ - $RT$ d) $PV$ - $RT$ / $d$ c) $PV$ - $RT$ / $d$ c) $PV$ - $RT$ / $d$ d) $PV$ - $RT$ / $d$ when an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) $2/5$ b) $3/5$ c) $3/7$ d) $3/4$ 5. At zero Kelvin, which of the following properties of a gas will be zero? A 1 a) Kinetic energy b) Potential energy c) Mass d) Density  6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at $0$ C only c) $0$ Make it melt at $0$ S only $0$ C only c) $0$ A $0$ Density  7. If $1101$ of heat is supplied to a gaseous system, its internal energy changes by $0$ D. The amount of external work done is a) $0$ 150U b) $0$ 70U c) $0$ 110U d) $0$ 40U  8. A body cools from $0$ 50.0°C to $0$ 49.9°C in $0$ 5s. How long will it take to cool from $0$ 10°C to $0$ 39.9°C? Assume the temperature of the surroundings to be $0$ 30.0°C and Newton's law of cooling to be valid. a) $0$ 2.5s b) $0$ 10s c) $0$ 20s		-/			
a) 6400J  2. For hydrogen gas C <sub>P</sub> -C <sub>V</sub> =a and for oxygen gas C <sub>P</sub> -C <sub>V</sub> =b. The relation between a and b is a) a=16b b) a=b/16 c) a=4b d) a=b  3. The equation of state corresponding to 8 g of O₂ is (assume O₂ to be an ideal gas) a) PV=RTT b) PV=RTT d) PV=RT		-,			
2. For hydrogen gas $C_P$ - $C_v$ = $a$ and for oxygen gas $C_P$ - $C_v$ = $b$ . The relation between $a$ and $b$ is a $a$ = $16b$ b) $a$ = $b$ / $16$ c) $a$ = $a$ + $b$ d) $a$ = $b$ 3. The equation of state corresponding to $8$ g of $O_2$ is (assume $O_2$ to be an ideal gas)  a) $PV$ = $RT$ b) $PV$ = $RT$ c) $PV$ = $RT$ d) $PV$ = $RT$ / $2$ 4. When an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) $2/5$ b) $3/5$ c) $3/7$ d) $3/4$ 5. At zero Kelvin, which of the following properties of a gas will be zero?  a) Kinetic energy b) Potential energy c) Mass d) Density Density 6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at $0$ faster rate d) Raise its melting point c) Make it melt at $0$ faster rate d) Raise its melting point 7. If $1101$ of heat is supplied to a gaseous system, its internal energy changes by 40J. The amount of external work done is a) $150J$ b) $70J$ c) $110J$ d) $40J$ 8. A body cools from $50.0^{\circ}$ C to $49.9^{\circ}$ C in $5$ s. How long will it take to cool from $40.0^{\circ}$ C to $39.9^{\circ}$ C? Assume the temperature of the surroundings to be $30.0^{\circ}$ C and Newton's law of cooling to be valid. a) $2.5$ s b) $10$ s c) $20$ s		c) 3200J			
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b) a=b/16 c) a=4b d) a=b  3. The equation of state corresponding to 8 g of O <sub>2</sub> is (assume O <sub>2</sub> to be an ideal gas) a) PV=8RT b) PV=RT/4 c) PV=RT d) PV=RT/2  4. When an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) 2/5 b) 3/5 c) 3/7 d) 3/4  5. At zero Kelvin, which of the following properties of a gas will be zero? a) Kinetic energy b) Potential energy c) Mass d) Density  6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at a faster rate d) Raise its melting point c) Make it melt at a faster rate d) Raise its melting point 7. If 110J of heat is supplied to a gaseous system, its internal energy changes by 40J. The amount of external work done is a) 150J b) 70J c) 110J d) 40J  8. A body cools from 50.0°C to 49.9°C in 5s. How long will it take to cool from 40.0°C to 39.9°C? Assume the temperature of the surroundings to be 30.0°C and Newton's law of cooling to be valid. a) 2.5s b) 10s c) 20s	2.	7 0 0 7	D	1	
c) $a=4b$ d) $a=b$ 3. The equation of state corresponding to 8 g of $O_2$ is (assume $O_2$ to be an ideal gas)  a) $PV=8RT$ b) $PV=RT/4$ c) $PV=RT/2$ 4. When an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) $2/5$ b) $3/5$ c) $3/7$ d) $3/4$ 5. At zero Kelvin, which of the following properties of a gas will be zero? a) Kinetic energy b) Potential energy c) Mass d) Density  6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at $0$ colly c) Make it melt at $0$ colly c) Make it melt at a faster rate d) Raise its melting point 1f 110J of heat is supplied to a gaseous system, its internal energy changes by 40J. The amount of external work done is a) $150J$ b) $70J$ c) $110J$ d) $40J$ 8. A body cools from $50.0^{\circ}$ C to $49.9^{\circ}$ C in $5s$ . How long will it take to cool from $40.0^{\circ}$ C to $39.9^{\circ}$ C? Assume the temperature of the surroundings to be $30.0^{\circ}$ C and Newton's law of cooling to be valid. a) $2.5s$ b) $10s$ c) $20s$		-,			
d) a=b  3. The equation of state corresponding to 8 g of O2 is (assume O2 to be an ideal gas) a) PV=8RT b) PV=RT/4 c) PV=RT/4 c) PV=RT/2  4. When an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) 2/5 b) 3/5 c) 3/7 d) 3/4  5. At zero Kelvin, which of the following properties of a gas will be zero? a) Kinetic energy b) Potential energy c) Mass d) Density  6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at 0°C only c) Make it melt at a faster rate d) Raise its melting point 7. If 110J of heat is supplied to a gaseous system, its internal energy changes by 40J. The amount of external work done is a) 150J b) 70J c) 110J d) 40J  8. A body cools from 50.0°C to 49.9°C in 5s. How long will it take to cool from 40.0°C to 39.9°C? Assume the temperature of the surroundings to be 30.0°C and Newton's law of cooling to be valid. a) 2.5s b) 10s c) 20s		-/ " " " " "			
3. The equation of state corresponding to 8 g of O <sub>2</sub> is (assume O <sub>2</sub> to be an ideal gas) a) PV=8RT b) PV=RT d) PV=RT d) PV=RT d) PV=RT/2  4. When an ideal monoatomic gas is heated at constant pressure, the fraction of heat energy supplied which increases the internal energy of the gas is a) 2/5 b) 3/5 c) 3/7 d) 3/4  5. At zero Kelvin, which of the following properties of a gas will be zero? A 1 a) Kinetic energy b) Potential energy c) Mass d) Density  6. By exerting a certain amount of pressure on an ice block, you a) Lower its melting point b) Make it melt at 0°C only c) Make it melt at a faster rate d) Raise its melting point 1f 110J of heat is supplied to a gaseous system, its internal energy changes by 40J. The amount of external work done is a) 150J b) 70J c) 110J d) 40J  8. A body cools from 50.0°C to 49.9°C in 5s. How long will it take to cool from 40.0°C to 39.9°C? Assume the temperature of the surroundings to be 30.0°C and Newton's law of cooling to be valid. a) 2.5s b) 10s c) 20s					
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b) 10s c) 20s					
		c) 20s			
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S. No	Questions	Answer	Marks	Chapter
9.	The critical temperature of CO <sub>2</sub> is 31.1°C and the room temperature is 40°c, then CO <sub>2</sub> behaves as a  a) Gas  b) Vapour  c) Gas and vapour  d) Liquid	A	1	
10.	One mole of a monoatomic gas is mixed with one mole of a diatomic gas. What will be the value of $\gamma$ for the mixture?  (a) 1.5  (b) 1.54  (c) 1.4  (d) 1.45	A	1	
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