

# Rajiv Gandhi University of Health Sciences, Karnataka

## I Year B.Pharm Degree Examination – 13-Jan-2020

**Time: Three Hours****Max. Marks: 70 Marks**

### **PHARMACEUTICAL INORGANIC CHEMISTRY** **(Revised Scheme 3)** **Q.P. CODE: 2605**

Your answers should be specific to the questions asked.  
Draw neat labeled diagrams wherever necessary.

**LONG ESSAYS (Answer any Two)****2 x 10 = 20 Marks**

1. How impurities are incorporated in pharmaceuticals? Discuss the importance of limit tests in quality control of Pharmaceuticals.
2. What are redox titrations? Classify them with examples. Define reducing agents and oxidizing agents with examples. Write the assay principle of Copper sulphate and Hydrogen peroxide.
3. Define precipitation titration. Explain the different methods of Precipitation titrations in detail.

**SHORT ESSAYS (Answer any Six)****6 x 5 = 30 Marks**

4. What are primary, secondary standard substances? Give examples Give the standardization of 0.1 N Potassium permanganate.
5. How do you carry out the limit test for sulphate in the given sample of sodium bicarbonate and sodium benzoate?
6. What are pharmaceutical aids? Classify them with suitable examples.
7. What are antacids? Classify them with examples Give the ideal properties of antacids.
8. Define Inhalant? Explain the methods of preparation, uses and labeling conditions of Oxygen and Carbon dioxide.
9. Give a method of preparation, assay principle, medicinal use of magnesium sulphate with its chemical formula and synonym if any.
10. Write the preparation and standardization of 0.1N Ferric ammonium sulphate.
11. Write a note on iron and ammonium citrate.

**SHORT ANSWERS****10 x 2 = 20 Marks**

12. Define acidimetric and alkali metric titrations.
13. What are astringents? Give example.
14. Significance figures.
15.  $\text{NH}_4\text{Cl} + \text{HCHO} + \text{H}_2\text{O} \rightarrow$
16. Role of acetic acid and ammonia in the limit test for heavy metals.
17. Define emetics with examples.
18. Give the medicinal use of sodium thiosulphate and sodium meta bi sulphate.
19. Why sulphuric acid is added in the assay of Hydrogen per oxide.
20. Role of potassium thiocyanate in the assay of copper sulphate.
21. Classify nonaqueous solvents.

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