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B.Tech I Year I Semester (R15) Supplementary Examinations June 2018 ENGINEERING CHEMISTRY

(Common to ECE, ME, EIE & IT)

Max. Marks: 70

Time: 3 hours

1

PART – A

(Compulsory Question)

- Answer the following: (10 X 02 = 20 Marks)
 - (a) Using simple chemical reactions, explain how water can be disinfected by chlorination.
 - (b) What is meant by DO with reference to water chemistry?
 - (c) Give two examples each of thermoplastic and thermoset polymers.
 - (d) What are living polymers?
 - (e) What is an electrochemical cell?
 - (f) Define passivity of a metal.
 - (g) Write the main steps involved in the refining of crude oil.
 - (h) Define calorific value of a fuel.
 - (i) Write the general composition of Portland cement.
 - (j) What are refractories? Mention the important characteristics of refractory materials.

PART – B

(Answer all five units, 5 X 10 = 50 Marks)

UNIT – I

2 What is meant by softening of water? Explain the process of water softening using ion-exchange resins. What other ion-exchangers can be used for this purpose? Which of these are more advantages.

OR

3 What are boiler troubles? Explain in short caustic embrittlement and its preventive measures.

4 Discuss the preparation, properties and applications of Nylon-6, Nylon-6, 6 and Nylon – 11.

OR

5 Define elastemers. Briefly explain how is natural rubber processed, vulcanized and compounded.

UNIT – III

6 Derive an equation for calculating electrode potential of an electrode. State the importance of finding electrode potentials and also state what parameters affect electrode potential of an electrode.

OR

7 Using one example, explain what is concentration cell corrosion. How can we prevent concentration cell corrosion?

UNIT – IV

8 Describe Otto-Hoffmann's by product over process of manufacturing metallurgical coke.

OR

9 Describe the Bergius process of making synthetic petrol.

UNIT – V

10 Explain using a neat diagram, the thin film theory of lubricants.

OR

11 What are CNTs? How are they prepared? Where do they find applications?

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