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11. HYDROGEN

$$\left\{ \frac{111 \text{ At. Wt. of H}}{100} \times \frac{1.1 \times 14}{100} \right\} \% \text{ F} \times 3$$

$$\left\{ \frac{2 \text{ Vapour density of a gas}}{\text{molar mass of H}_2} \right\} \times 100$$

LHS and RHS both are unitless.

Molecular wt. =

$$\left\{ \frac{24 \text{ Vol. Strength of H}_2}{\text{Normality} \times T_{1,2}} \right\}$$

$$= \text{Normality} \times 5,6$$