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ORGANIC CHEMISTRY

Relations for the estimation of elements in organic compounds_

$$\% \text{ C} = \frac{12}{44} \times \frac{\text{mass of CO}_2 \text{ formed}}{\text{Mass of the compound}} \times 100$$

$$\% \text{ of H} = \frac{2}{18} \times \frac{\text{Mass of H}_2\text{O formed}}{\text{Mass of the compound}} \times 100$$

$$\% \text{ N} = \frac{14}{22400} \times \frac{\text{Volume of N}_2 \text{ at NTP}}{\text{Mass of compound}} \times 100$$

$$\% \text{ of N} = \frac{\text{Normality of acid used} \times \text{Volume of acid used}}{\text{Mass of the compound}} \times 100$$

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	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{H} \end{array}$	$\text{CH}_3 - \text{CH}_2 - \text{CHO}$			
2.	Two different types of dissimilar H-atoms $\text{CH}_3 - \text{COCH}_2\text{CH}_3$	One type of similar H-atoms (0-14 - 042 - a-K)	3	3	6
3	Two different types of dissimilar H-atoms $\text{Cl-CH}_2\text{CO} - \text{CH}_2 - \text{CH}_3$	Two different types of dissimilar H-atoms $\text{Ph} - \text{CH}_2 - \text{CO} - \text{CH}_2 - \text{CH}_3$	4	4	8
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