STATE OF MATTER

10. Van der Waals Equation : $[P + \frac{n^2}{V}(V - nb) = nRT$

12 Combined Gas Equiatr-on : $\frac{F1\,Vi\ P2V,}{T_i} =$ Ft constant}

(3.11. Dalton's law of partial pressures : p . P_i p. P_i ,

(4) Ideal gas equation = nRT

M = d RT

(SI Cherie:es Law $\frac{VI}{T1}$ = — (at cc:instant P and II)

(E}. Avogadroili Law V = kn (at constant P and T)

(4 Partial Pressure in Terms of Mole Fraction : P_i

{811. Boyla's Law : (at constant' and n)

(Di, Corn pres5ibility factor Z $\frac{PV}{RT}$ (for 1 mole of Gas)