

Roll No.

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Total No. of Pages : 02

Total No. of Questions : 10

B.Pharma (2011 to 2016) (Sem.-2)

PHARMACEUTICAL CHEMISTRY-II (Physical Chemistry)

Subject Code : BPHM-202

Paper ID : [D1113]

Time : 3 Hrs.

Max. Marks : 80

INSTRUCTION TO CANDIDATES :

1. SECTION-A is COMPULSORY consisting of FIFTEEN questions carrying TWO marks each.
2. SECTION-B contains FIVE questions carrying FIVE marks each and students have to attempt any FOUR questions.
3. SECTION-C contains FOUR questions carrying TEN marks each and students have to attempt any THREE questions.

SECTION-A**1. Explain in brief :**

- a. Ideal gas equation.
- b. Define and give example of Additive properties.
- c. Temperature coefficient.
- d. Partition Coefficient and its significance.
- e. Colligative properties.
- f. Ideal and Non ideal Solution.
- g. What do you mean by the term Paracahor?
- h. Conductance and its unit.
- i. Differentiate between heterogeneous and homogeneous catalysis.
- j. Boyles law.
- k. Define and give suitable example of positive catalysis.

1. Schrodinger wave equation.
- m. Isotherms.
- n. Expression for Pusedo first order reactions.
- o. State variables.

SECTION-B

2. Define and give the formula of Viscosity and discuss its applications in structure elucidation.
3. Lineweaver Burk Plot.
4. Draw a neat clean labelled Jablenski Digram.
5. Discuss the limitations of first Law of Thermodynamics.
6. Derive the expression for acid base catalytic reactions.

SECTION-C

7. Discuss the reasons of ideal gas deviation and derive the expression for Vander Waal's Gas equation.
8. Discuss in detail the two different theories of catalysis explaining behaviour of enzymatic reactions.
9. Define and derive the expression for Beer lamber law. Discuss its applications.
10. What is Phase rule and explain different type of equilibrium in connection with phase rule?